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must be committed to memory.

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Note: Not all of the reactions will occur.
For those that do not, write no reaction.

1. A piece of copper is dropped into a
container of water. No reaction
2. Liquid bromine is added to a container of

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sodium iodide crystals. $\text{Br (l)} + 2\text{NaI(s)}$
 $2\text{NaBr(s)} + \text{I}_2\text{(s)}$ 3. An aluminum strip is
immersed in a solution of silver nitrate.

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Practice-Pg. 49 • Formation of a Gas -
Pg. 50, Practice-Pg. 51 • Acid Base
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do not break apart and are in net ionic
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by Hague and Smith 7 Free elements, no
matter how complex the molecule, have
an oxidation number (valence or charge)
equal to zero OJ Fe T~ He < /1 1 ~ ~ L-td
8 The following are diatomic or
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Handbook SnCO tin(II) carbonate sodium
hydrogen carbonate NaHCO 1. 2. 3. 4. 5.
6. 8. 9. 10. 11. vanadium(V) oxide H₂O
dihydrogen monoxide (NH)₄CO
ammonium oxalate 4 224 polonium(VI)
thiocyanate Po(SCN)₆ tetraphosphorus
decaoxide P₄O₁₀ 12. 13. 15. 16. 17.
18. 19. manganese(VII) oxide copper(II)
dihydrogen phosphate

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The Ultimate Chemical Equations Handbook Determine the oxidation number of the underlined element:
 KMnO_4 Since K is an alkali metal, its charge must be $1+$ Oxygen is $2-$ but there are four of them, therefore, 4 times $2-$ equals $8-$ If $1+$ and $8-$ are

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